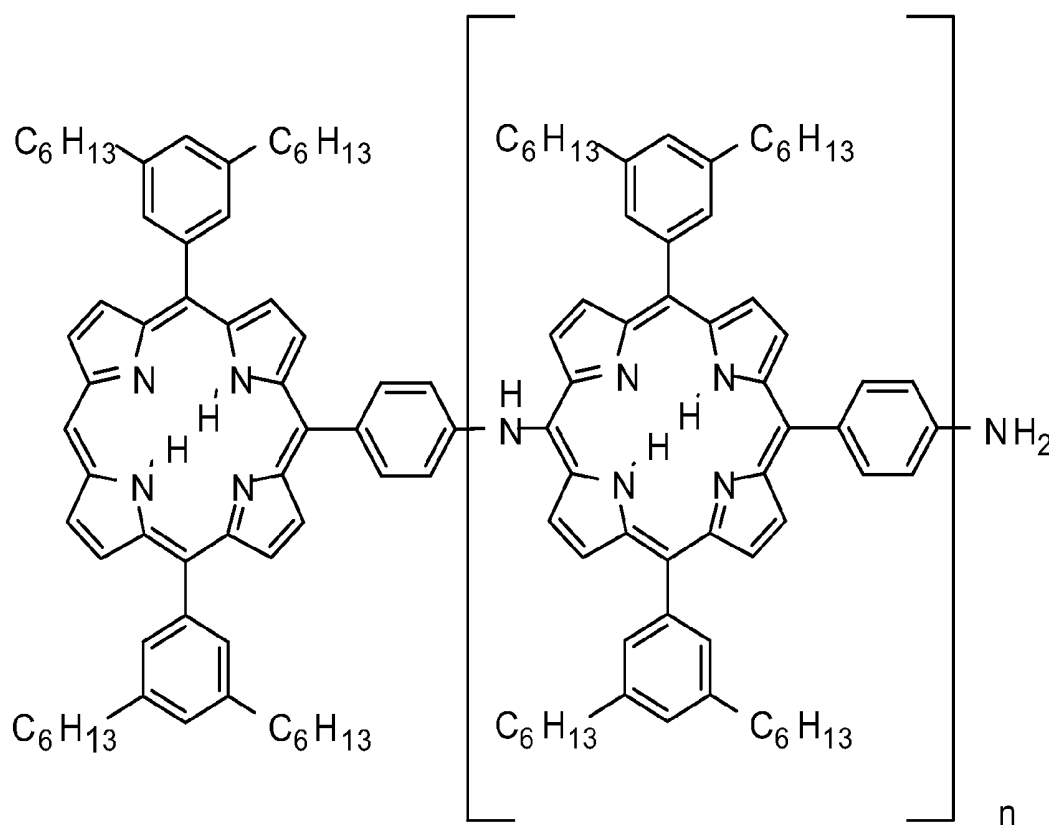


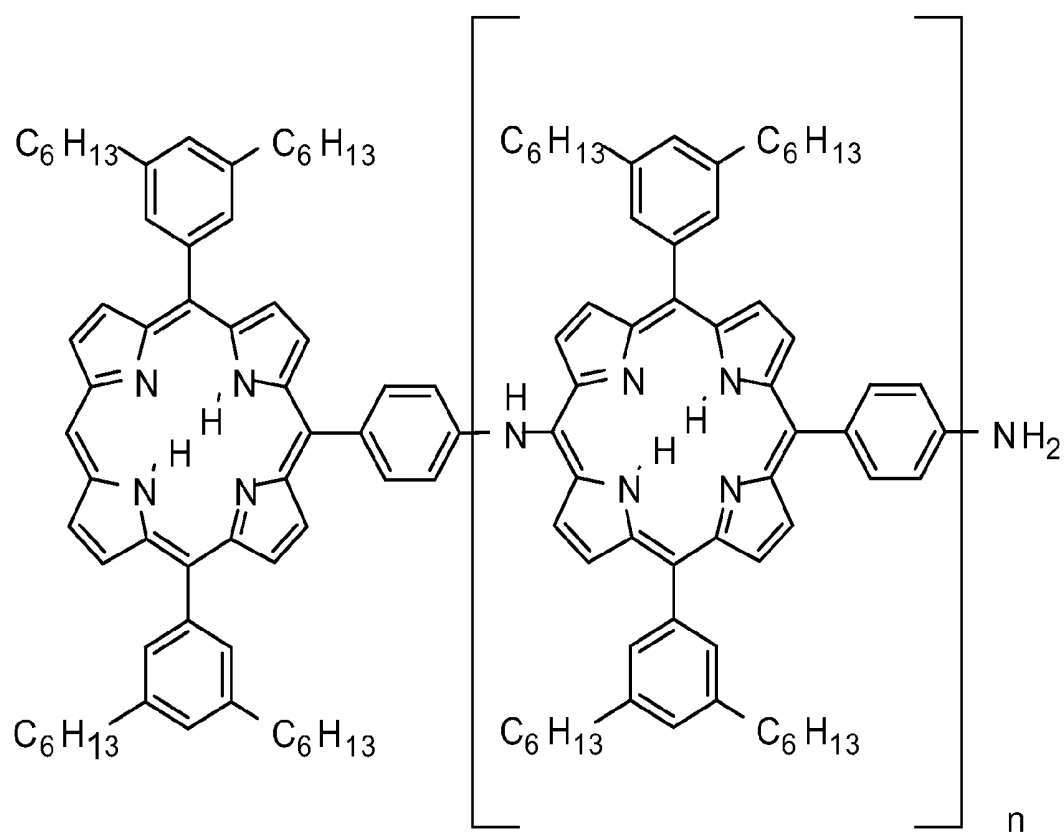


US 20130324686A1

(19) **United States**(12) **Patent Application Publication**
Gust, JR. et al.(10) **Pub. No.: US 2013/0324686 A1**(43) **Pub. Date: Dec. 5, 2013**(54) **SOLUBLE PORPHYRIN POLYMERS****Publication Classification**(71) Applicants: **John Devens Gust, JR.**, Mesa, AZ (US);
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AZ (US)(51) **Int. Cl.**
H01L 51/00 (2006.01)(72) Inventors: **John Devens Gust, JR.**, Mesa, AZ (US);
Paul Anthony Liddell, Apache Junction,
AZ (US)(52) **U.S. Cl.**
CPC **H01L 51/0077** (2013.01); **H01L 51/0035**
(2013.01)USPC **526/259**; 540/145(21) Appl. No.: **13/907,452**(22) Filed: **May 31, 2013****Related U.S. Application Data**(60) Provisional application No. 61/654,463, filed on Jun.
1, 2012.(57) **ABSTRACT**

Porphyrin polymers of Structure 1, where n is an integer (e.g., 1, 2, 3, 4, 5, or greater) are synthesized by the method shown in FIGS. 2A and 2B. The porphyrin polymers of Structure 1 are soluble in organic solvents such as 2-MeTHF and the like, and can be synthesized in bulk (i.e., in processes other than electropolymerization). These porphyrin polymers have long excited state lifetimes, making the material suitable as an organic semiconductor for organic electronic devices including transistors and memories, as well as solar cells, sensors, light-emitting devices, and other opto-electronic devices.

**Structure 1**



Structure 1

FIG. 1

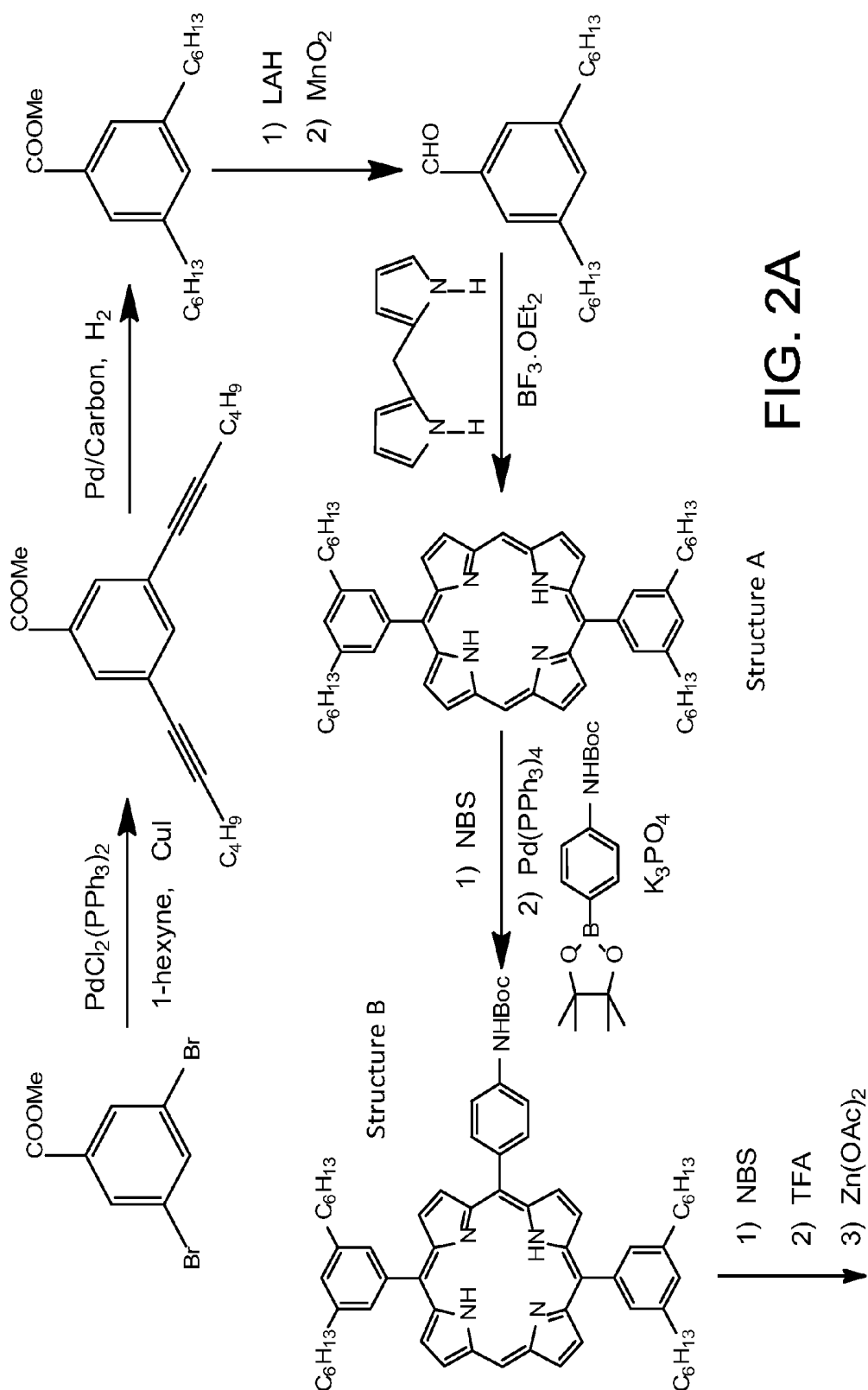


FIG. 2A

Structure A

Structure B

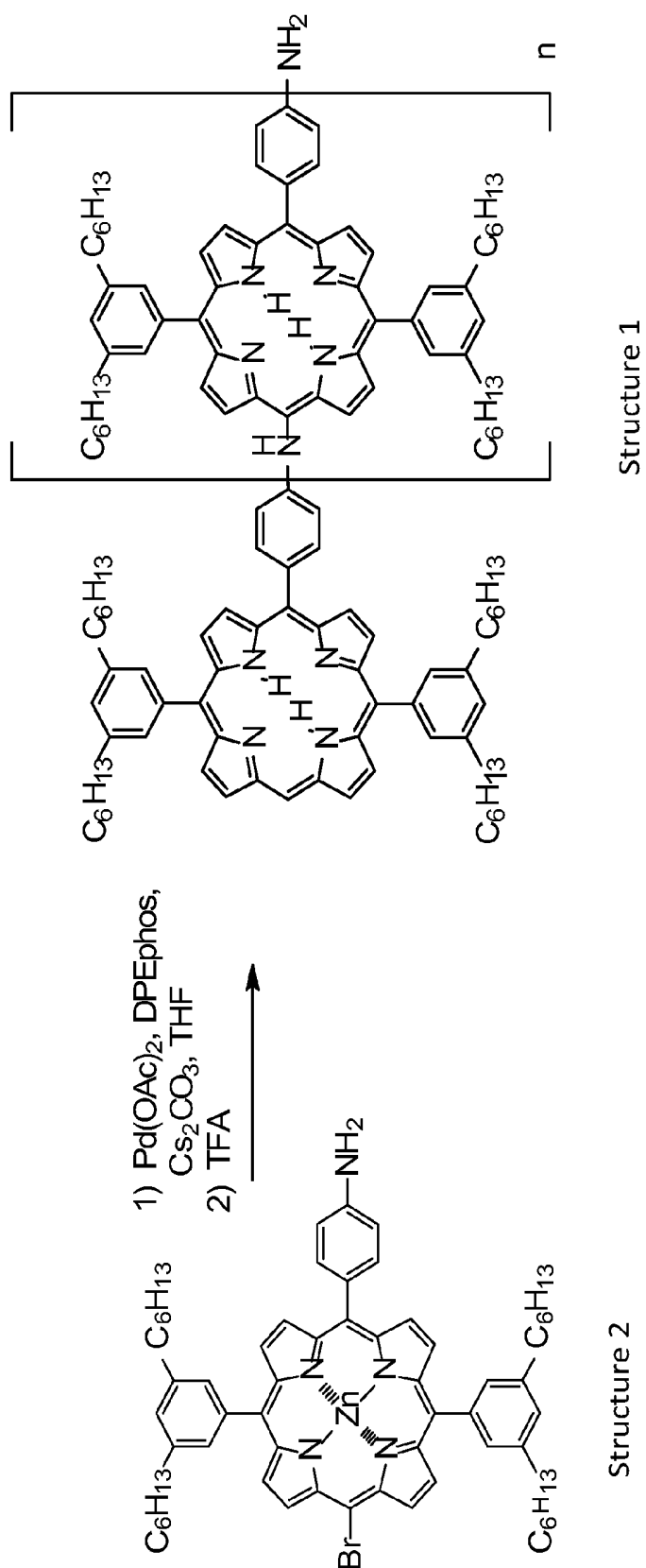


FIG. 2B

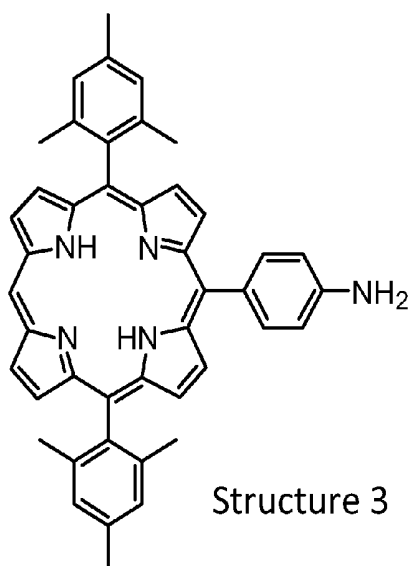


FIG. 3

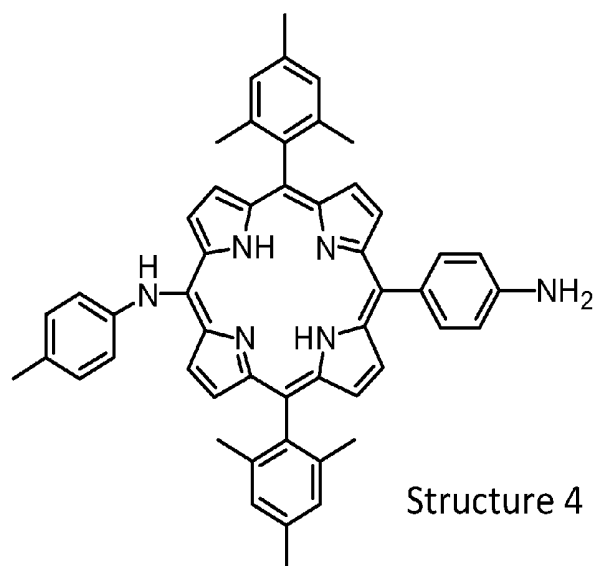


FIG. 4

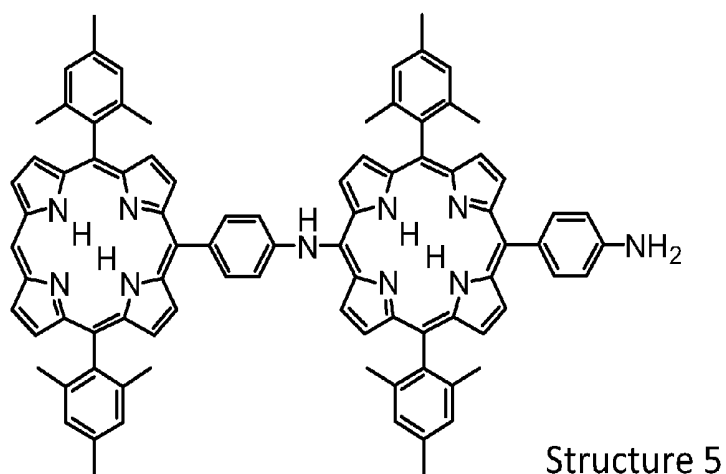


FIG. 5

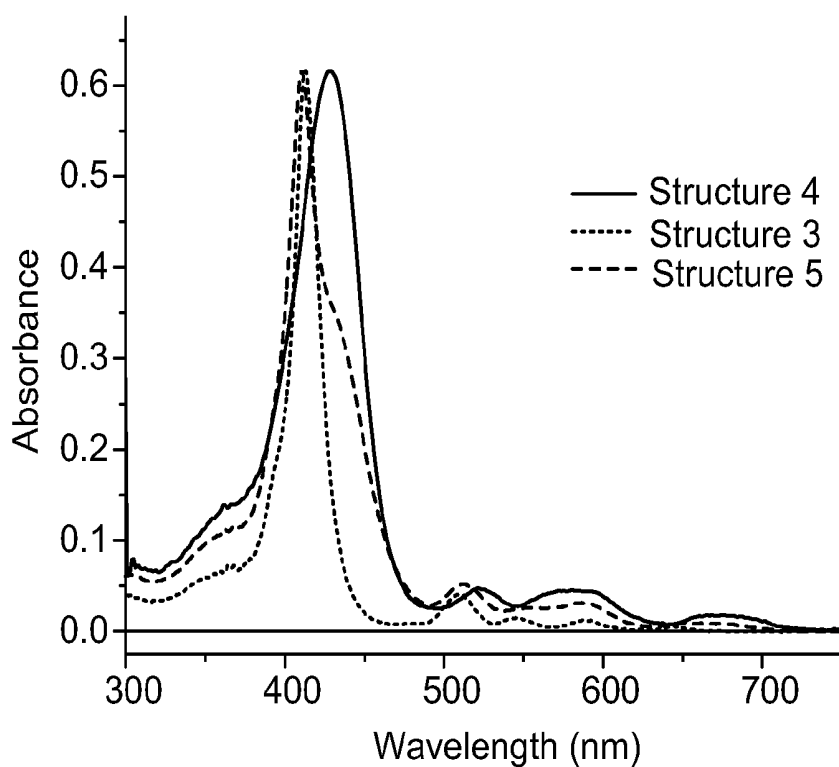


FIG. 6

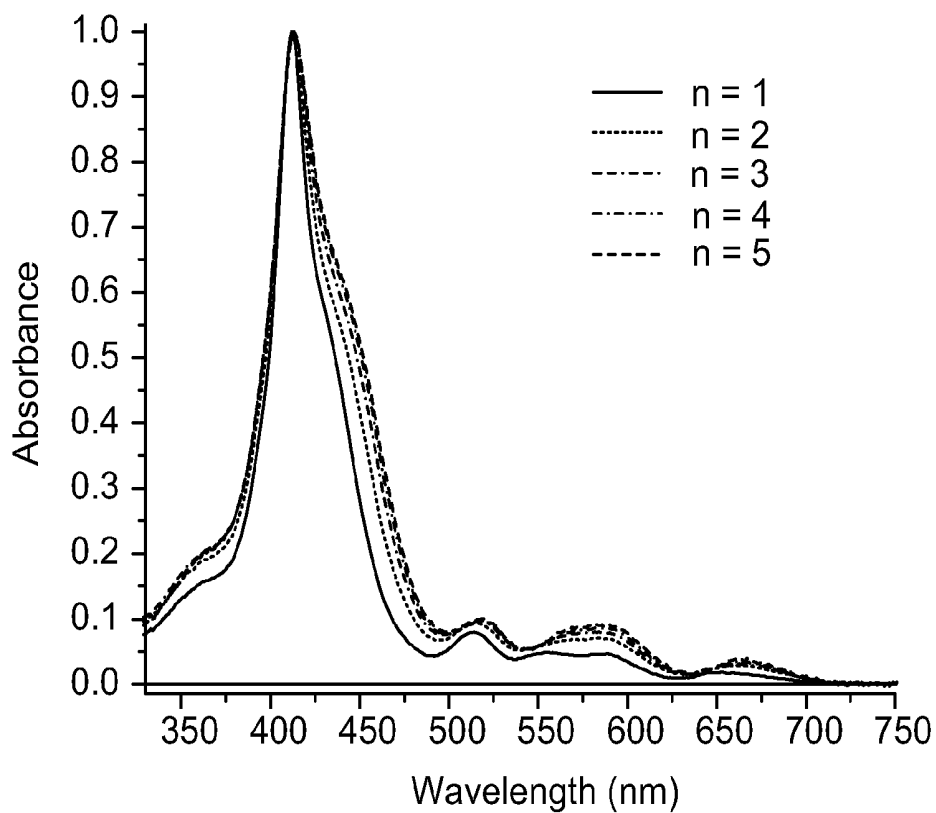
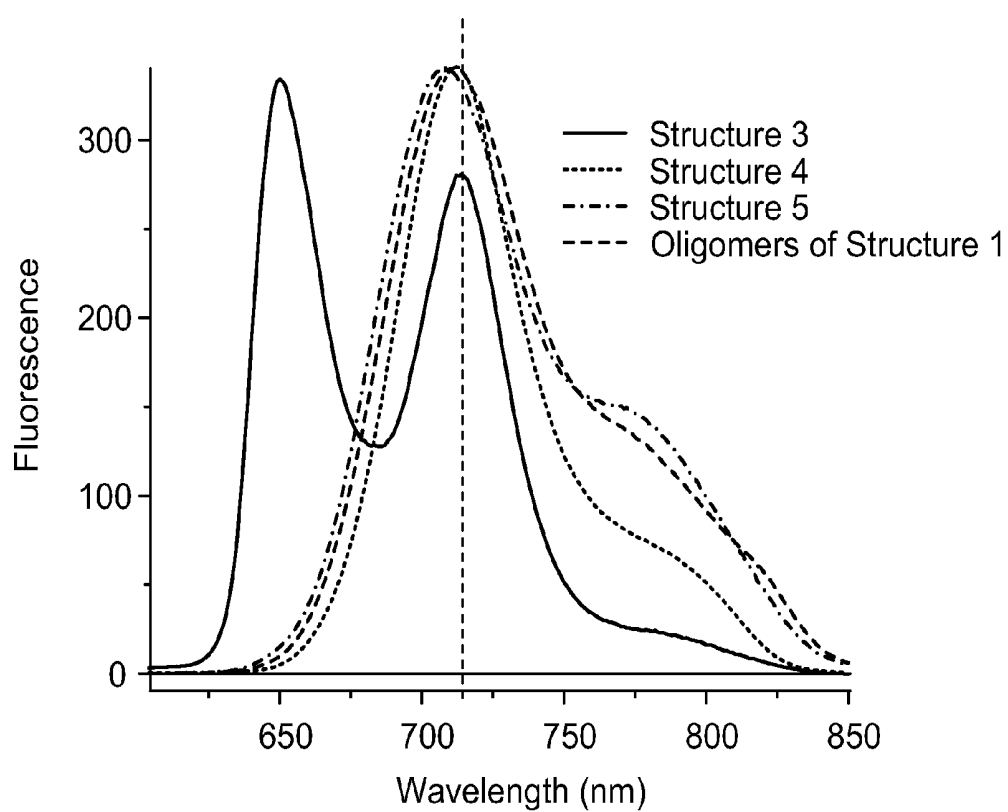


FIG. 7

**FIG. 8**

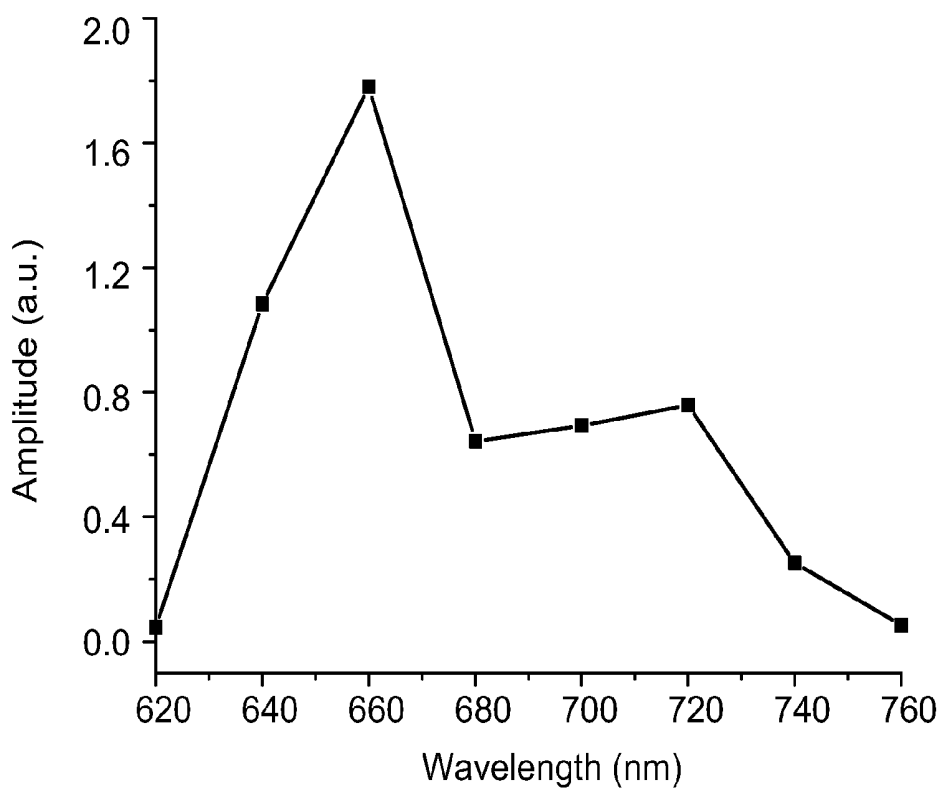


FIG. 9

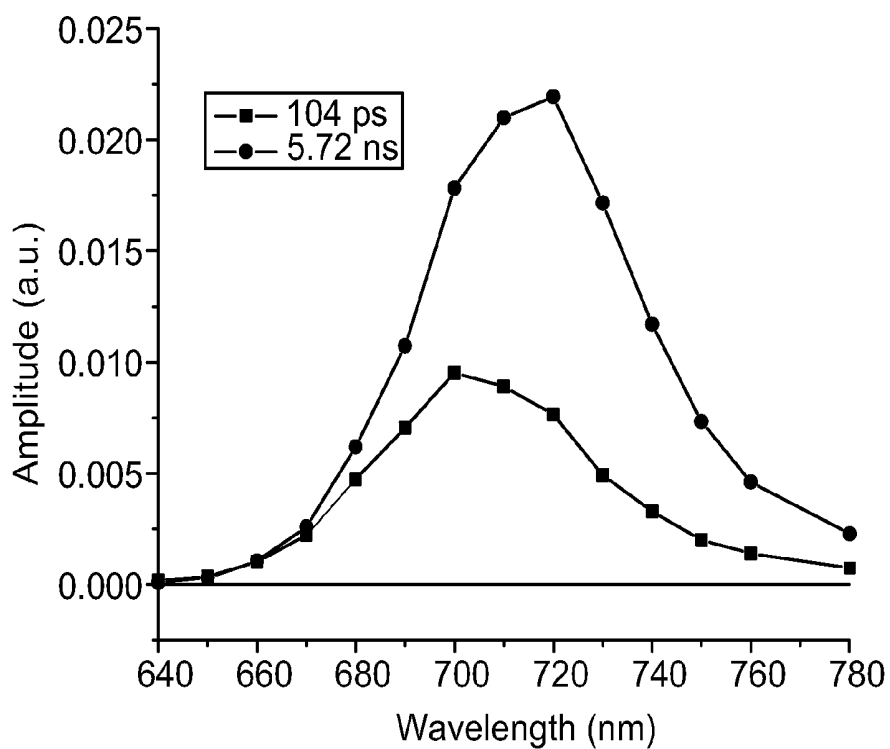


FIG. 10

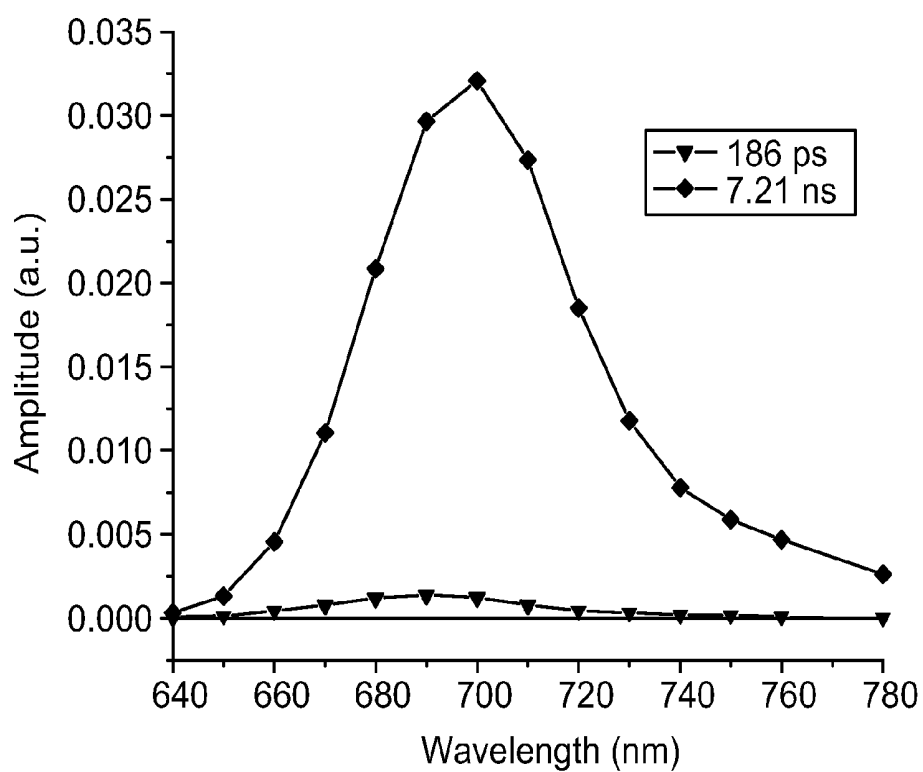


FIG. 11

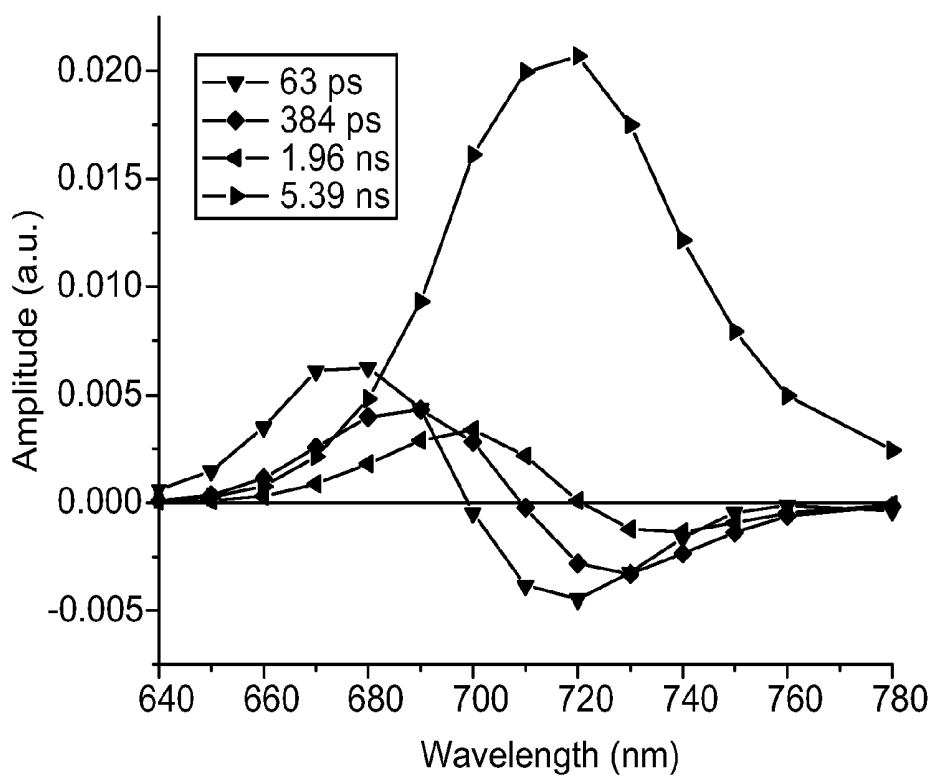


FIG. 12

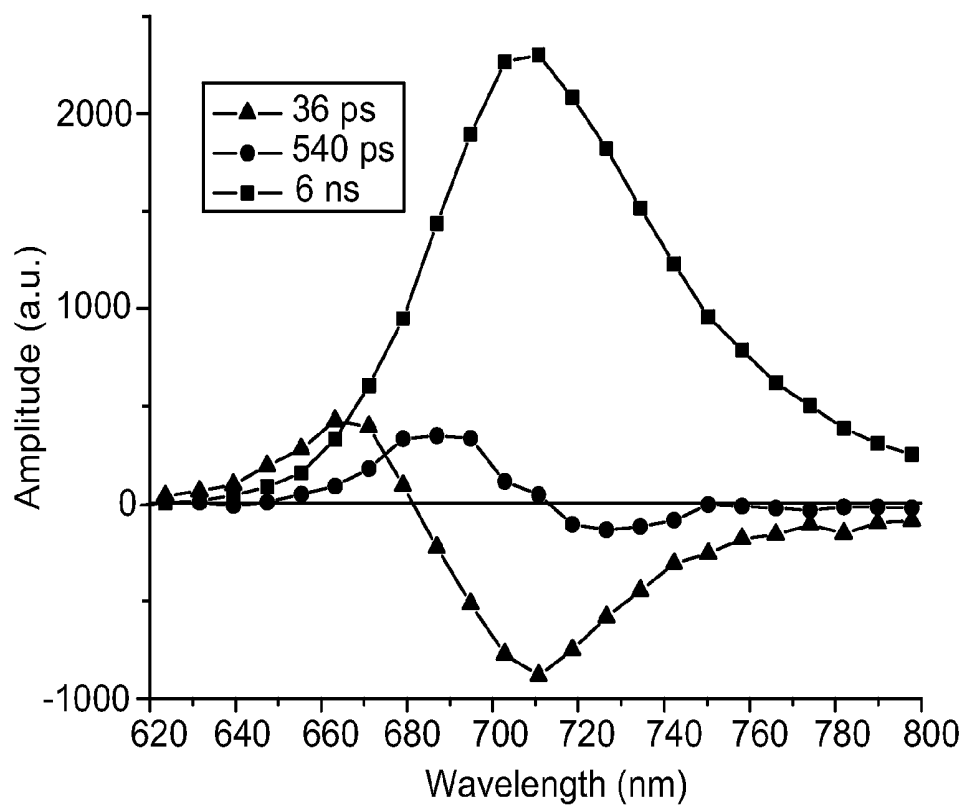


FIG. 13

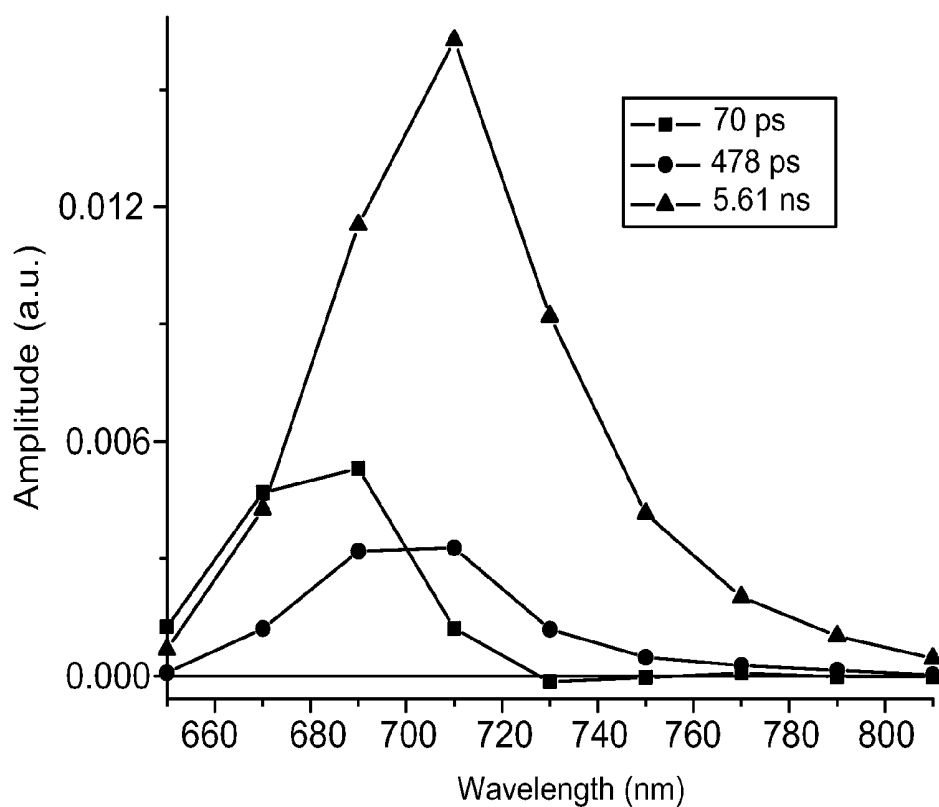
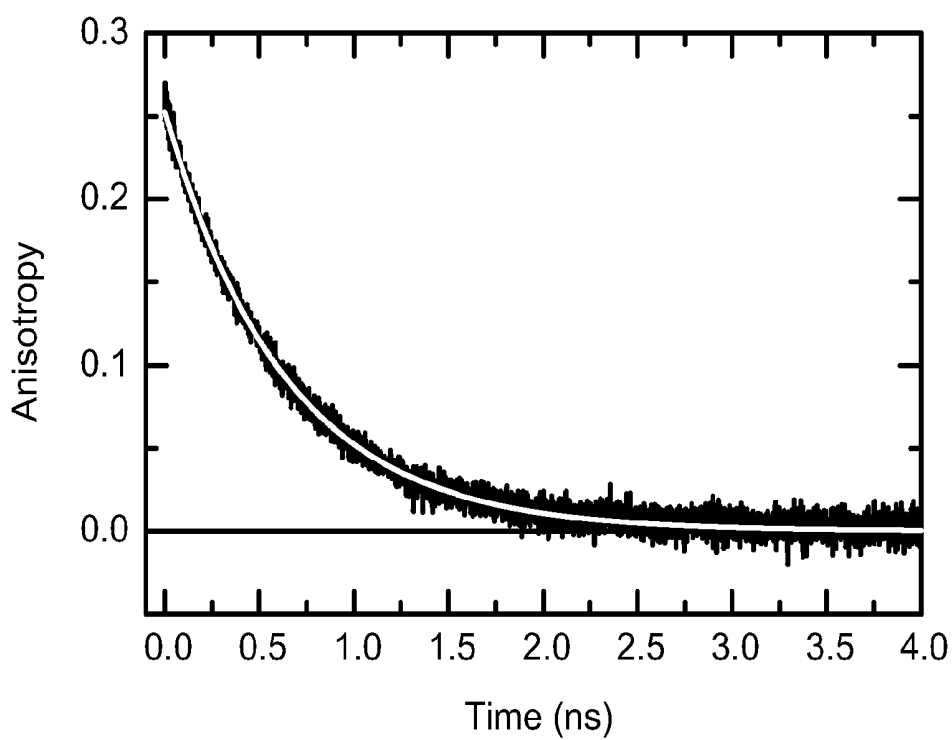
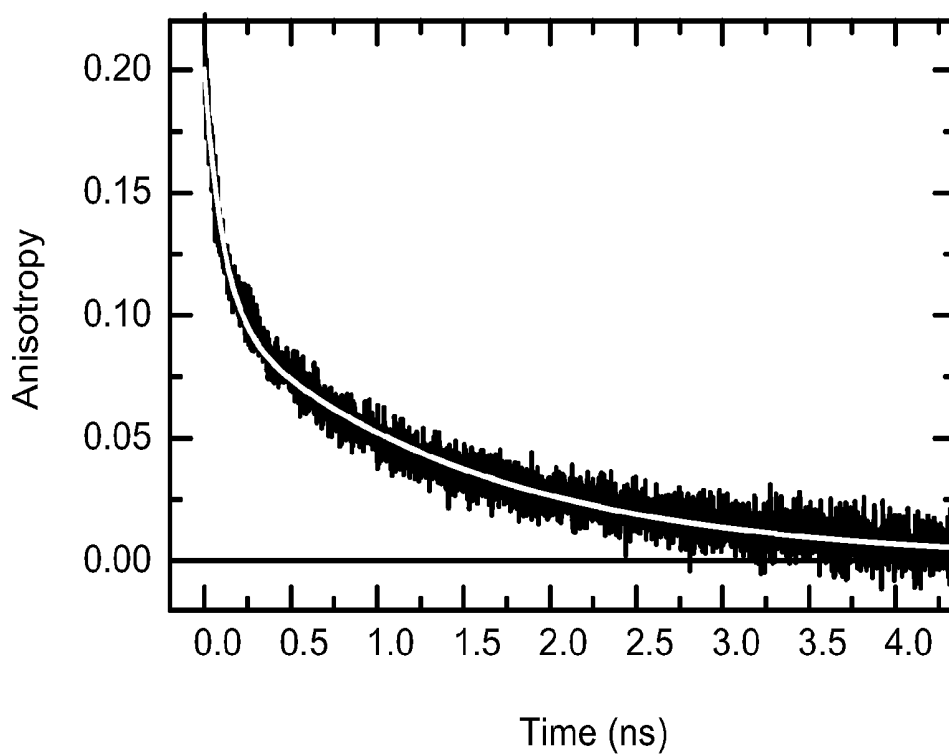
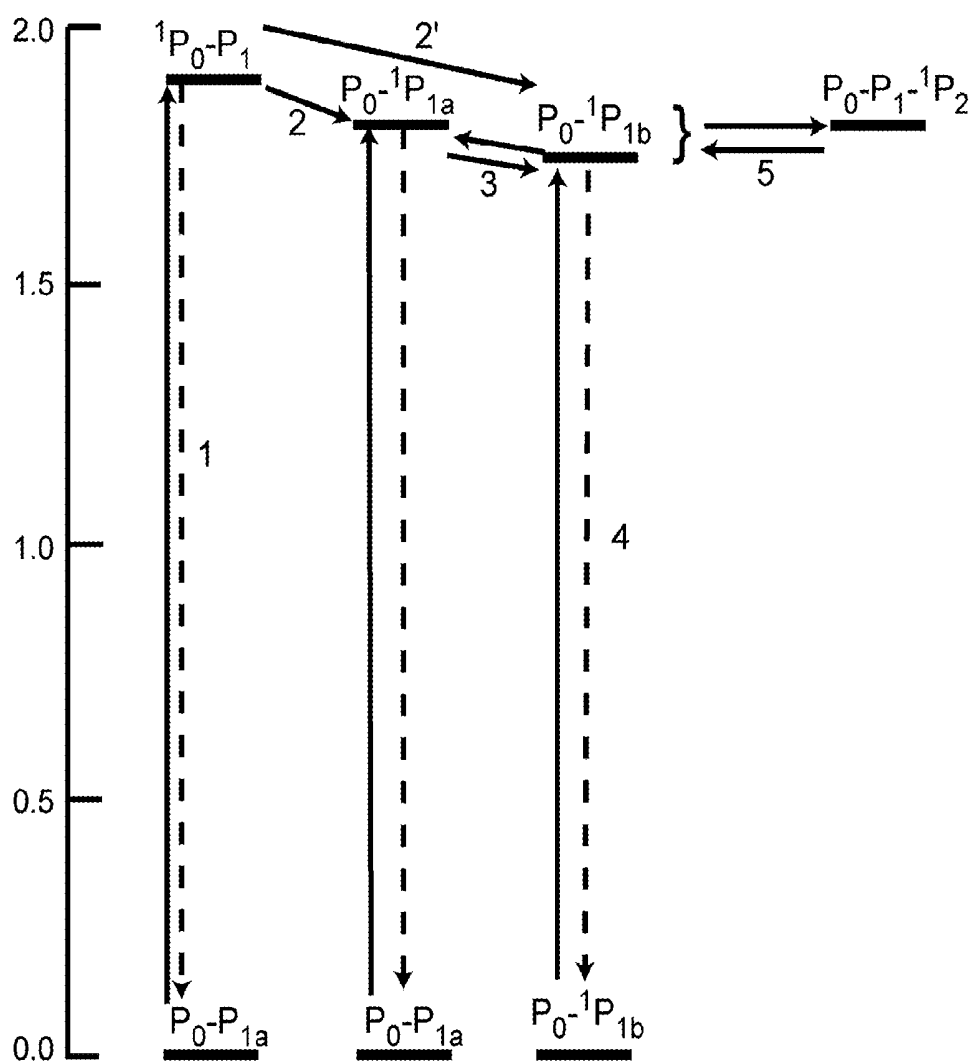


FIG. 14

**FIG. 15****FIG. 16**



$$\tau_1 = 9.5 \text{ ns}$$

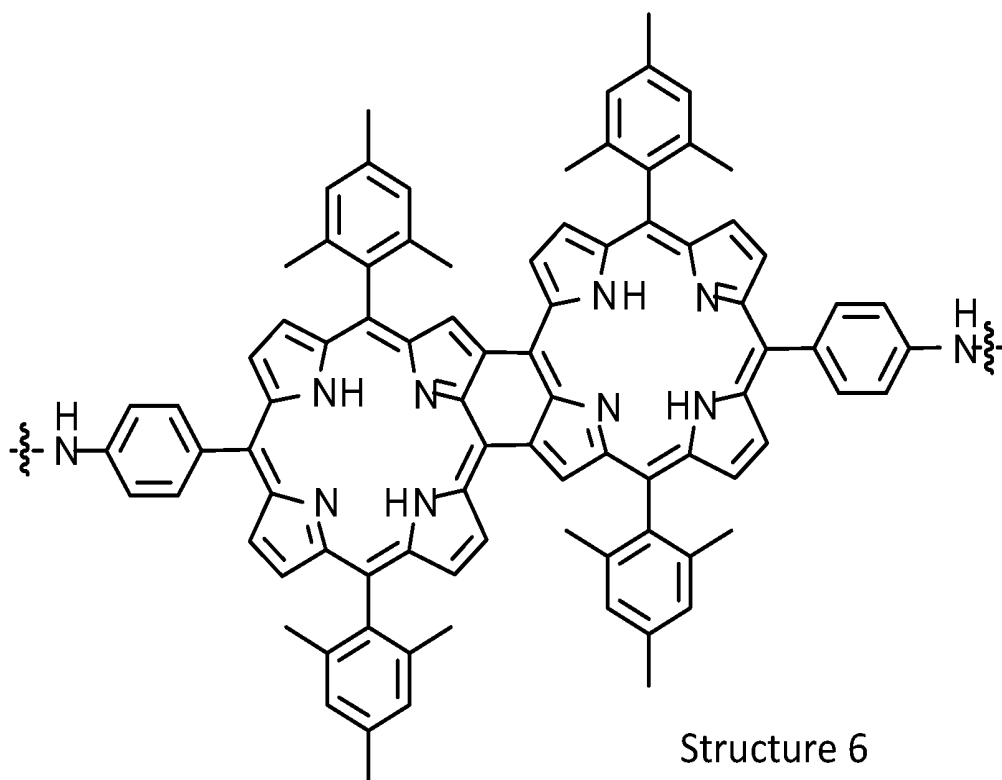
$$\tau_2 \text{ and } \tau_{2'} \ll 9.5 \text{ ns}$$

$$\tau_3 = \sim 500 \text{ ps}$$

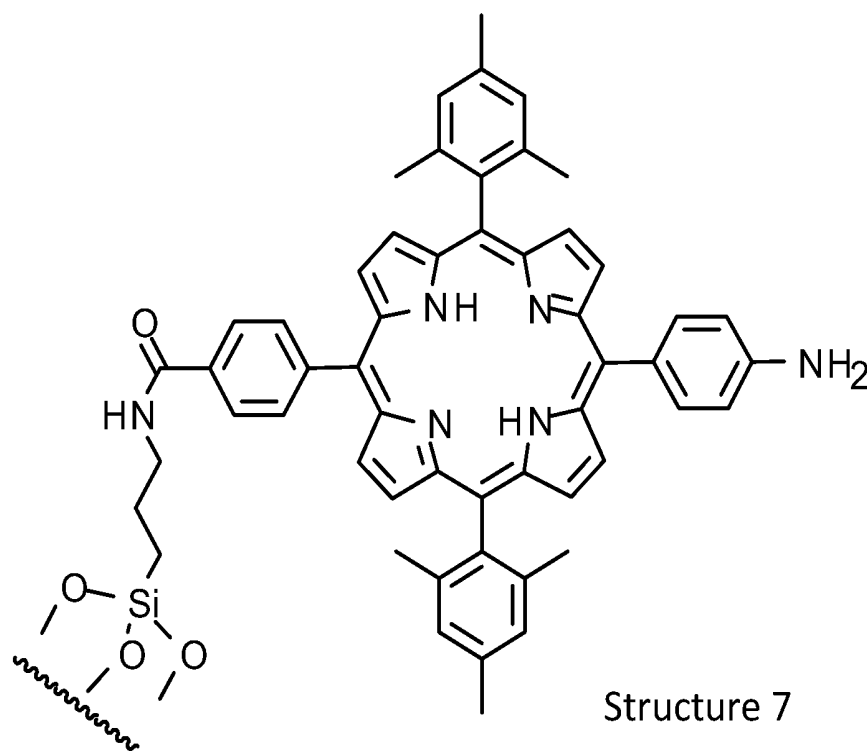
$$\tau_4 = \sim 6 \text{ ns}$$

$$\tau_5 = \tau_{-5} = 210 \text{ ps}$$

FIG. 17



Structure 6
FIG. 18



Structure 7

FIG. 19

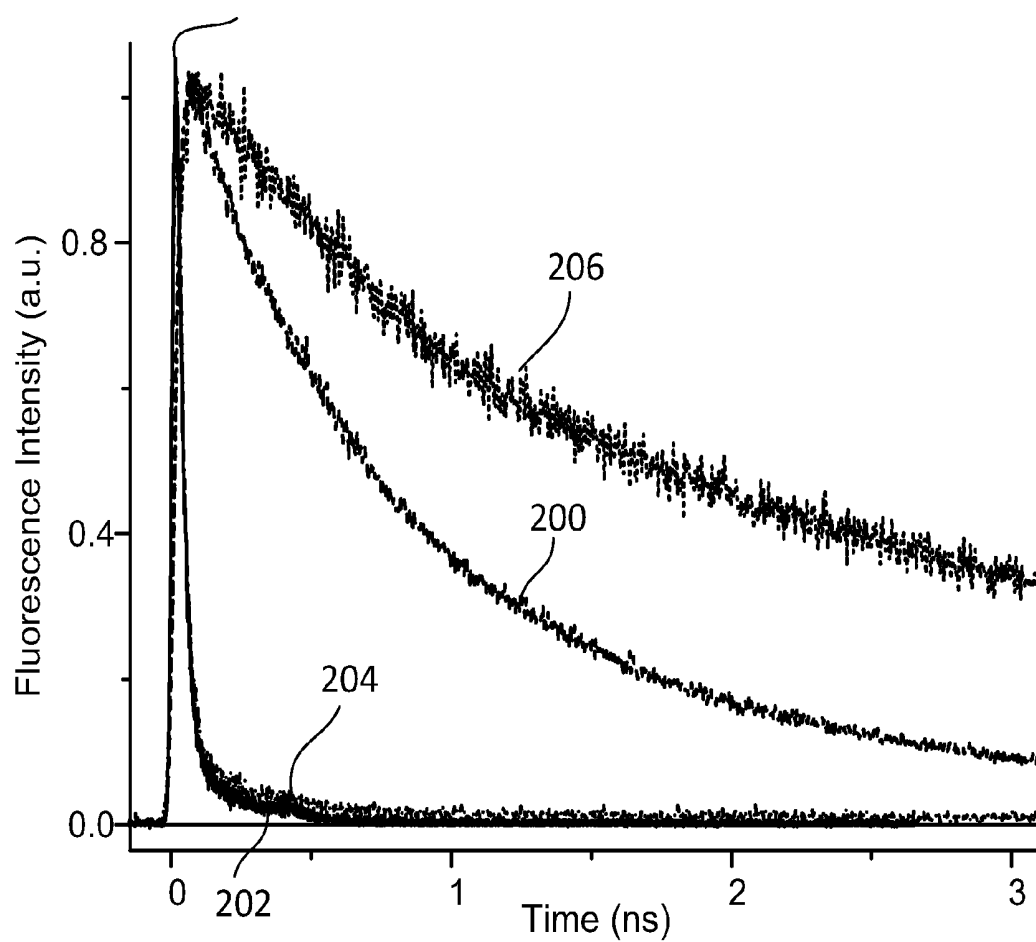


FIG. 20

SOLUBLE PORPHYRIN POLYMERS**CROSS-REFERENCE TO RELATED APPLICATION**

[0001] This application claims priority to U.S. Provisional Application Ser. No. 61/654,463 filed on Jun. 1, 2012, which is incorporated herein by reference in its entirety.

STATEMENT OF GOVERNMENT INTEREST

[0002] This invention was made with government support under DE-FG02-03ER15393 awarded by the Department of Energy. The government has certain rights in the invention.

TECHNICAL FIELD

[0003] This invention relates to chemical synthesis of conducting, soluble, solution-processable porphyrin polymers.

BACKGROUND

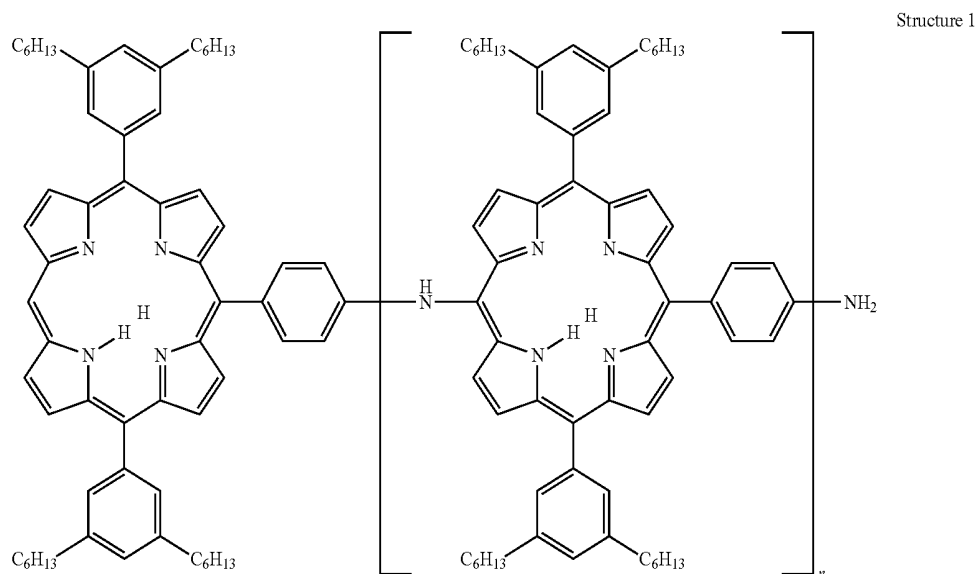
[0004] One of the major applications of conducting polymers is as a component of an “organic” or “bulk heterojunction” photovoltaic cell for making electricity from sunlight. Organic solar cells generally include a layer of transparent

organic substituent groups attached to the basic polymer backbone. The substituents can change the electrical properties of the polymer, its processing and mechanical properties, and its compatibility/interfacing with other materials. Typically, it is possible to write delocalized structures for these polymers that allow conductivity through the polymer backbone. Often, some type of “doping” is required to achieve good conductivity.

[0006] Such organic photovoltaic cells have very low efficiencies (i.e., less than about 5%) due to incomplete conversion of excitons to charge separation, recombination of electrons and holes within the organic layer or at the electrodes, and a lack of efficient light absorption throughout the solar spectral range. The lack of electron acceptors covalently bonded to the polymer strands can lead to inefficient conversion of excitons to charge separation, rapid charge recombination, high concentrations of carriers at interface areas (leading to increased recombination), the need to prepare bicontinuous phases of polymer and fullerene or other electron carriers (which adds to processing requirements), and the need to use excessive amounts of the electron transport phase.

SUMMARY

[0007] Porphyrin polymers of Structure 1, where n is an integer (e.g., 1, 2, 3, 4, 5, or greater)



indium tin oxide (ITO) on glass, a layer of hole-conducting polymer mixed with an organic or inorganic electron-conducting phase, and a third layer of a low-work-function metal such as aluminum. Light is absorbed by the polymer in the photovoltaic cell; the resulting excitation then migrates to a phase boundary with the electron-conducting phase, and an electron is injected from the polymer. The electrons migrate to the metal through the electron-conducting phase, and holes migrate to the ITO via the polymer, generating a photocurrent.

[0005] Conducting polymers for use in organic light-emitting diodes (OLEDs) and organic conductors (“wires”), some of which may be electropolymerized, have been reported. These polymers include, for example, polyacetylenes, polyphenyleneethynylenes (PPE), polyphenylenevinylenes (PPV), polythiophenes, and polyanilines. The structural variations of these polymers are typically changes in the

are synthesized by the method shown in FIGS. 2A and 2B. The porphyrin polymers of Structure 1 are soluble in organic solvents such as 2-methyltetrahydrofuran (2-MeTHF) and the like, and can be synthesized in bulk (i.e., in processes other than electropolymerization). These porphyrin polymers are conducting (e.g., semiconducting) and have long excited state lifetimes, making the material suitable as an organic semiconductor for organic electronic devices including transistors and memories, as well as solar cells, sensors, light-emitting devices, and other opto-electronic devices.

BRIEF DESCRIPTION OF THE DRAWINGS

[0008] FIG. 1 depicts a porphyrin polymer having Structure 1.

[0009] FIGS. 2A and 2B depict synthesis of porphyrin polymers having Structure 1.

[0010] FIG. 3 depicts a model porphyrin monomer having Structure 3.

[0011] FIG. 4 depicts a model porphyrin monomer having Structure 4.

[0012] FIG. 5 depicts a porphyrin dyad having Structure 5.

[0013] FIG. 6 shows absorbance spectra of porphyrin polymers having Structures 3-5 in 2-methyltetrahydrofuran (2-MeTHF).

[0014] FIG. 7 shows normalized absorbance spectra of polymers having Structure 1.

[0015] FIG. 8 shows normalized emission spectra of polymers having Structures 3-5 and a mixture of oligomers having Structure 1 in 2-MeTHF.

[0016] FIG. 9 shows a decay-associated spectrum of emission from a porphyrin polymer having Structure 3 in 2-MeTHF, excitation at 420 nm.

[0017] FIG. 10 shows decay-associated spectra of porphyrin polymers having Structure 4 in 2-MeTHF, excitation at 520 nm.

[0018] FIG. 11 shows decay-associated spectra of porphyrin polymers having Structure 4 in cyclohexane, excitation at 520 nm.

[0019] FIG. 12 shows decay-associated spectra of porphyrin polymers having Structure 4 in decanol, excitation at 520 nm.

[0020] FIG. 13 shows decay-associated spectra of porphyrin polymers having Structure 5 in 2-MeTHF, excitation at 520 nm.

[0021] FIG. 14 shows decay-associated spectra of porphyrin polymers having Structure 1 with $n=5$ in 2-MeTHF, excitation at 420 nm.

[0022] FIG. 15 shows fluorescence anisotropy of a porphyrin dyad having Structure 1 with $n=1$.

[0023] FIG. 16 shows fluorescence anisotropy of a porphyrin oligomer having Structure 1 with $n=3$.

[0024] FIG. 17 depicts a schematic diagram of transient states of Structure 1 and some of their interconversion pathways and time constants.

[0025] FIG. 18 depicts a porphyrin electropolymer having Structure 6.

[0026] FIG. 19 depicts a structure of a porphyrin polymer having Structure 7 bound to a surface.

[0027] FIG. 20 depicts normalized fluorescence decays of porphyrin electropolymers on indium tin oxide (ITO).

DETAILED DESCRIPTION

[0028] Films of porphyrin and porphyrin-fullerene dyad conducting polymers have been prepared as films on electrodes by electropolymerization, as described in Liddell et al., *Chem. Mater.* 2008, 20, 135-142; Gervald et al., *Photochem. Photobiol. Sci.* 2010, 9, 890-900; and U.S. 2010/0065123 entitled "Electrically Conducting Porphyrin and Porphyrin-Fullerene Electropolymers," all of which are incorporated herein by reference. However, use of these materials is limited by the insolubility of the films on the electrodes.

[0029] As described herein, solution, bulk chemical synthesis of a class of conducting (e.g., semiconducting), soluble, porphyrin polymers with long excited state lifetimes has been developed. This process yields a porphyrin polymer free or almost free of defects due to reverse linking of monomers, and thus free or almost free of excited state quenching and charge carrier trapping caused by such defects. Accordingly, the porphyrin polymer has long excited state lifetimes, making the material suitable as an organic semiconductor for organic electronic devices including transistors and memories, as well as solar cells, sensors, light-emitting devices, and other opto-electronic devices. Unlike earlier polymers pro-

duced by electropolymerization, these soluble polymers, based on a porphyrin monomer designed for polymer solubility, are solution processable, suitable for spin coating, and otherwise able to be handled and used as bulk materials.

[0030] FIG. 1 depicts Structure 1, a conducting porphyrin polymer, where n is an integer. The multiple hexyl chains on each porphyrin render the polymer soluble in organic solvents, or contribute at least in part to such solubility. Steady-state and transient spectroscopic studies have revealed some of the unusual photophysical properties of this material. FIGS. 2A and 2B depict synthesis of Structure 1 via Structure 2, a brominated porphyrin monomer. FIGS. 3 and 4 depict porphyrin monomers having Structures 3 and 4, respectively, and FIG. 5 depicts Structure 5, a porphyrin dyad.

[0031] FIG. 6 shows absorbance spectra of Structures 3, 4, and 5 in 2-methyltetrahydrofuran (2-MeTHF). FIG. 7 shows normalized absorbance spectra of Structure 1 with $n=1-5$. FIG. 8 shows normalized emission spectra of Structures 3 and 4, dyad 5, and a mixture of oligomers of Structure 1 in 2-MeTHF. In the dyad and oligomers, only emission from the porphyrins bearing nitrogen at a meso-position is seen due, at least in part, to rapid energy transfer from the porphyrin with H at the meso-position.

[0032] Time-resolved emission studies were carried out by the single-photon-timing method in 2-MeTHF. FIG. 9 shows a decay-associated spectrum (DAS) of emission from Structure 3 in 2-MeTHF with excitation at 420 nm ($\tau=9.53$ ns, $\chi^2=1.06$). FIG. 10 shows DAS of Structure 4 in 2-MeTHF with excitation at 520 nm. Two components with time constants of 104 ps and 5.72 ns ($\chi^2=1.16$) were observed. FIG. 11 shows DAS of Structure 4 in cyclohexane, excitation at 520 nm. Two components with time constants of 186 ps (plot 110) and 7.21 ns (plot 112) were observed. FIG. 12 shows DAS of Structure 4 in decanol, excitation at 520 nm. FIG. 13 shows DAS of Structure 5 in 2-MeTHF, excitation at 520 nm. FIG. 14 shows DAS of Structure 1, $n=5$, in 2-MeTHF, excitation at 420 nm. Roughly comparable DAS spectra were observed for Structure 4 and polymers of Structure 1 of various lengths. DAS similar to the above were obtained for Structure 5 with excitation at 450 nm, where the unit without a meso nitrogen does not absorb. Thus, none of the components appear to represent energy transfer.

[0033] Fluorescence anisotropy experiments were carried out on Structure 1 ($n=1$) and on Structure 1 ($n=3$) with $\lambda_{ex}=680$ nm and $\lambda_{em}=780$ nm. FIG. 15 shows Structure 1, $n=1$, and FIG. 16 shows Structure 1, $n=3$. The results show a time constant for excitation hopping between porphyrins in the polymer chain of 210 ps.

[0034] Structure 1 is readily soluble in organic solvents known in the art due at least in part to the hexyl chains. The absorption spectrum is broad, and extends to wavelengths longer than 700 nm. There is a weak electronic interaction between porphyrin moieties observed in absorption. The absorption and emission spectra of Structure 4 and the polymers of Structure 1 are solvent polarity dependent, suggesting an excited state with charge transfer character. Photochemical properties of the polymers of Structure 1 are summarized in FIG. 17, which shows a schematic diagram of transient states of polymers of Structure 1 and some of their interconversion pathways and time constants. The polymer of Structure 1 is represented as $P_0-P_1-P_n$. Excitation of P_0 is followed by very rapid energy transfer to P_1 , which is also formed by direct absorption. P_1 exists in several conformations. The initially formed excited states of P_1 relax through two or more short-lived species ($\sim 50-600$ ps in 2-MeTHF), likely involving different conformations about the C-N bonds, to the least-energetic excited state. The 1P_1 state is fluorescent and decays

in ~6 ns, but can also undergo singlet energy transfer to yield $^1P_2 \dots ^1P_n$ with a time constant of 210 ps. In some solvents, additional excited state conformations are present.

[0035] Molecules similar to Structure 2 but lacking the bromine electropolymerize, but electropolymerization can introduce defects such as Structure 6, shown in FIG. 18, that act as excitation and charge traps. FIG. 19 shows ITO-bound Structure 7. FIG. 20 shows normalized fluorescence decays of electropolymers on ITO, including the polymer of Structure 6 on ITO-bound Structure 7 (plot 200), the polymer of Structure 6 on ITO (plot 202), the polymer of Structure 6 on silane-treated ITO (plot 204), and Structure 7 bound to ITO via an aminopropylsiloxy group (plot 206). Polymers such as those shown in Structure 6 have very short excited state lifetimes, as shown in FIG. 20. When the electrode surface is first derivatized with a porphyrin of Structure 7, formation of such traps cannot occur, or do not occur widely. Electropolymeric films prepared in this way show much longer excited state lifetimes with charge transfer character and energy migration between porphyrins, and thus are suitable for construction of solar cells or other devices.

SYNTHESIS

[0036] Porphyrin polymers having Structure 1 are synthesized by the method shown in FIGS. 2A and 2B as described below.

[0037] 5,15-Bis(3,5-Dihexylphenyl)Porphyrin (Structure A).

[0038] To a flask containing 3.3 g (22.59 mmol) of 2,2'-dipyrrromethane, 6.20 g (22.68 mmol) of 3,5-dihexylbenzaldehyde and 2.3 L of chloroform was added 1.72 mL of boron trifluoride diethyl etherate. After stirring under an argon atmosphere in the dark for 30 min, 2,3-dichloro-5,6-dicyano-1,4-benzoquinone (DDQ) (5.14 g) was added to the reaction mixture and stirring was continued for another 1 h. The dark solution was reduced in volume to approximately 1 L by evaporation at reduced pressure and then gently shaken with aqueous sodium bicarbonate (1 L). Once the two layers had separated the organic phase was washed a further three times with fresh bicarbonate solution. The organic layer was then concentrated to a viscous oil and this material was chromatographed on silica gel (hexanes/methylene chloride, 5:1 to 3:1). The appropriate fractions were combined, the solvent was evaporated at reduced pressure, and the residue was recrystallized from methylene chloride/methanol to give 3.564 g (40% yield) of the desired porphyrin. ^1H NMR (400 MHz) δ -3.08(2H, s, —NH), 0.92-0.95(12H, t, J=7 Hz, —CH₃), 1.37-1.45(16H, m, —CH₂—), 1.50-1.58(8H, m, —CH₂—), 1.92-1.85(8H, m, —CH₂—), 2.92(8H, t, J=8 Hz, —CH₂—), 7.43(2H, s, Ar-H), 7.91(4H, s, Ar-H), 9.12(4H, d, J=5 Hz, β —H), 9.38(4H, d, J=5 Hz, β —H), 10.30(2H, s, meso-H); MALDI-TOF-MS m/z calcd. for C₅₆H₇₀N₄ 798.6, obsd. 798.5; UV/vis (CH₂Cl₂) 408, 503, 537, 576, 631 (nm).

[0039] 5,15-Bis(3,5-Dihexylphenyl)-10-Bromoporphyrin.

[0040] To a flask containing 1.00 g (1.25 mmol) of 5,15-bis(3,5-dihexylphenyl)porphyrin and 200 mL of chloroform was added 223 mg (1.25 mmol) of N-bromosuccinimide (NBS). After stirring the reaction mixture for 15 min, TLC (hexanes/methylene chloride, 2:1) indicated that the product was present together with lesser amounts of the starting material and the dibrominated porphyrin. The reaction mixture was concentrated to a viscous oil and this material was chromatographed on silica gel (hexanes/methylene chloride, 4:1/3:1) to give 1.116 g (68% yield) of the desired product. ^1H NMR (400 MHz) δ -2.97(2H, s, —NH), 0.93(12H, t, J=7 Hz, —CH₃), 1.36-1.44(16H, m, —CH₂—), 1.48-1.56(8H, m, —CH₂—), 1.83-1.91(8H, m, —CH₂—), 2.89(8H, t, J=8 Hz, —

CH₂—), 7.43(2H, s, Ar-H), 7.85(4H, s, Ar-H), 8.99(2H, d, J=4 Hz, β —H), 9.00(2H, d, J=4 Hz, β —H), 9.28(2H, d, J=4 Hz, β —H), 9.73(2H, d, J=5 Hz, β —H), 10.16(1H, s, meso-H); MALDI-TOF-MS m/z calcd. for C₅₆H₆₉N₄Br 876.5, obsd. 876.4; UV/vis (CH₂Cl₂) 416, 512, 547, 588, 645 (nm).

[0041] 5,15-Bis(3,5-Dihexylphenyl)-10-(4-Tert-Butylphenylcarbamate)Porphyrin (Structure B).

[0042] To a heavy walled glass tube was added 1.00 g (1.14 mmol) of 5,15-bis(3,5-dihexylphenyl)-10-bromoporphyrin, 3.64 g (11.39 mmol) of 4-(Boc-amino)benzeneboronic acid pinacol, 4.83 g (22.78 mmol) of potassium phosphate tribasic and 20 mL of tetrahydrofuran (THF). The suspension was flushed with a stream of argon gas for 10 min, 132 mg (0.11 mmol) of tetrakis(triphenylphosphine) palladium(0) was added and the argon flushing procedure was continued for an additional 10 min. The tube was then sealed with a Teflon™ screw plug and warmed to 67° C. After 17 h, the tube was cooled and TLC (hexanes/methylene chloride, 1:1) of the contents indicated that all the starting material had been consumed. The reaction mixture was filtered through celite and the filtrate was concentrated to dryness. This material was chromatographed on silica gel (hexanes/methylene chloride, 2:1/1:1) to give 985 mg (87% yield) of the required porphyrin. ^1H NMR (400 MHz) δ -2.98(2H, s, —NH), 0.92(12H, t, J=7 Hz, —CH₃), 1.32-1.42(16H, m, —CH₂—), 1.46-1.56(8H, m, —CH₂—), 1.64(9H, s, —CH₃), 1.82-1.88(8H, m, —CH₂—), 2.88(8H, t, J=8 Hz, —CH₂—), 6.83(1H, s, —NH), 7.41(2H, s, Ar-H), 7.75(2H, d, J=8 Hz, Ar-H), 7.82(4H, s, Ar-H), 8.13(2H, d, J=8 Hz, Ar-H), 8.89(2H, d, J=4 Hz, β —H), 8.93(2H, d, J=4 Hz, β —H), 9.04(2H, d, J=4 Hz, β —H), 9.32(2H, d, J=4 Hz, β —H), 10.19(1H, s, meso-H); MALDI-TOF-MS m/z calcd. for C₆₇H₈₃N₅O₂ 989.6, obsd. 989.6; UV/vis (CH₂Cl₂) 414, 510, 545, 584, 639 (nm).

[0043] 5,15-Bis(3,5-Dihexylphenyl)-10-Bromo-20-(4-Tert-Butylphenylcarbamate)Porphyrin.

[0044] To a 1 L flask containing 2.90 g (2.93 mmol) of 5,15-bis(3,5-dihexylphenyl)-10-(4-tert-butylphenylcarbamate)porphyrin and 400 mL of chloroform was added 547 mg (3.07 mmol) of N-bromosuccinimide. The solution was stirred for 20 min. TLC (hexanes/methylene chloride, 1:1) indicated that the reaction was complete. The reaction mixture was washed with aqueous sodium bicarbonate, dried over anhydrous sodium sulfate, and then concentrated to dryness under reduced pressure. The resulting material was chromatographed on silica gel (hexanes/methylene chloride, 5:2/3:2) to give 2.813 g (90% yield) of the desired porphyrin. ^1H NMR (400 MHz) δ -2.74(2H, s, —NH), 0.91(12H, t, J=7 Hz, —CH₃), 1.34-1.41(16H, m, —CH₂—), 1.46-1.52(8H, m, —CH₂—), 1.64(9H, s, —CH₃), 1.81-1.88(8H, m, —CH₂—), 2.87(8H, t, J=8 Hz, —CH₂—), 6.82(1H, s, —NH), 7.41(2H, s, Ar-H), 7.75(2H, d, J=8 Hz, Ar-H), 7.82(4H, s, Ar-H), 8.10(2H, d, J=8 Hz, Ar-H), 8.82(4H, s, β —H), 8.92(2H, d, J=4 Hz, β —H), 9.65(2H, d, J=5 Hz, β —H); MALDI-TOF-MS m/z calcd. for C₆₇H₈₂N₅O₂Br 1067.6, obsd. 1067.5; UV/vis (CH₂Cl₂) 422, 520, 556, 597, 654 (nm).

[0045] 5-(4-Aminophenyl)-20-Bromo-15,20-Bis(3,5-Dihexylphenyl)Porphyrin.

[0046] To a flask containing a solution of 1.5 g (1.40 mmol) of 5,15-bis(3,5-dihexylphenyl)-10-bromo-20-(tert-butylphenylcarbamate)porphyrin dissolved in 40 mL of methylene chloride was added 60 mL of trifluoroacetic acid. The green solution was stirred under an argon atmosphere for 20 min, after which time TLC (hexanes/methylene chloride, 1/1) indicated that the reaction was complete. The reaction mixture was diluted with 200 mL of methylene chloride and washed with water (200 mL) several times. The organic layer was then washed with aqueous sodium bicarbonate, dried over

anhydrous sodium sulfate, and then concentrated to dryness. The residue was chromatographed on silica gel (methylene chloride/hexanes, 3:4/1:1) to give 1.238 g (91% yield) of the required porphyrin. ^1H NMR (400 MHz) δ -2.71(2H,s,-NH), 0.91(12H,t,J=7 Hz,-CH₃), 1.34-1.42(16H,m,-CH₂-), 1.46-1.52(8H,m,-CH₂-), 1.81-1.88(8H,m,-CH₂-), 2.86(8H,t,J=8 Hz,-CH₂-), 3.98(2H,s,-NH), 7.02(2H,d,J=8 Hz,Ar-H), 7.40(2H,s,Ar-H), 7.82(4H,s,Ar-H), 7.94(2H,d,J=8 Hz,Ar-H), 8.84(2H,d,J=5 Hz, β -H), 8.87(2H,d,J=5 Hz, β -H), 8.92(2H,d,J=5 Hz, β -H), 9.64(2H,d,J=5 Hz, β -H); MALDI-TOF-MS m/z calcd. for C₆₂H₇₄N₅Br 967.5, obsd. 967.5; UV/vis (CH₂Cl₂) 423, 521, 559, 598, 656 (nm). **[0047]** [5-(4-Aminophenyl)-15-Bromo-10,20-Bis(3,5-Dihexylphenyl)Porphyrino] Zinc(II) (Structure 2).

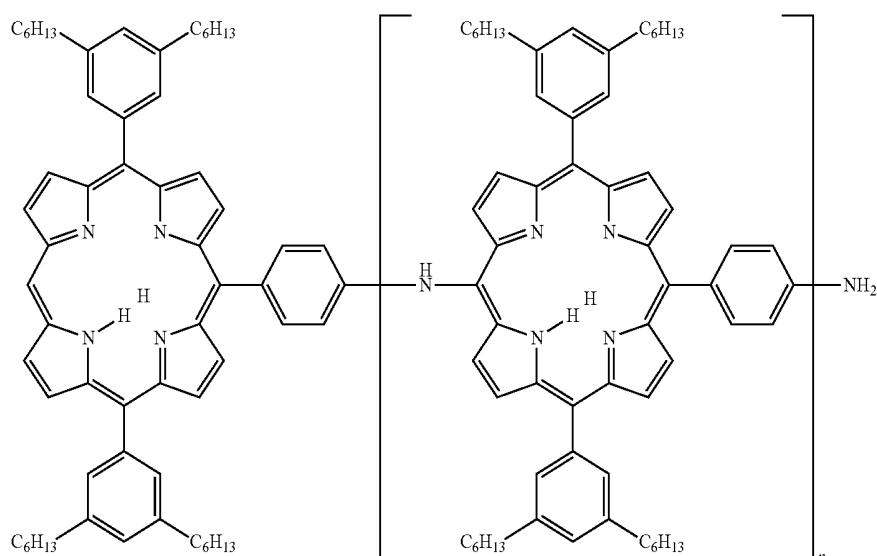
[0048] To a flask containing 1.20 g (1.24 mmol) of 5-(4-aminophenyl)-15-bromo-10,20-bis(3,5-dihexylphenyl)porphyrin and 200 mL of methylene chloride was added 50 mL of a saturated solution of zinc acetate dihydrate in methanol. After stirring for 1 h, TLC (hexanes/methylene chloride, 1/1) indicated that all the starting material had been consumed and a single product formed. The pink reaction mixture was washed with water (200 mL) several times and then with

was thoroughly washed with THF. The combined filtrate was concentrated and dried under high vacuum. This material was dissolved in a mixture of methylene chloride (200 mL) and trifluoroacetic acid (TFA) (200 mL) and stirred under an argon atmosphere for 1 h. The reaction mixture was then diluted with methylene chloride (200 mL) and then washed with water followed by aqueous sodium bicarbonate. The solvent was then evaporated and the residue was chromatographed on silica gel (methylene chloride/2% THF, methylene chloride/20% THF and finally methylene chloride/20% THF/10% methanol) to give three fractions of oligomeric material of Structure 1 ($n \geq 1$), the weight of each being, 0.43 g, 0.33 g, and 0.21 g.

[0051] A number of embodiments of the invention have been described. Nevertheless, it will be understood that various modifications may be made without departing from the spirit and scope of the invention. Accordingly, other embodiments are within the scope of the following claims.

What is claimed is:

1. A polymer having the structure:



aqueous sodium bicarbonate. It was then dried over anhydrous sodium sulfate. The solvent was evaporated and the remaining solid was dried under high vacuum to give 1.211 g (95% yield) of the metalated porphyrin. MALDI-TOF-MS m/z calcd. for C₆₂H₇₂N₅BrZn 1029.4, obsd. 1029.4; UV/vis (CH₂Cl₂) 424, 517(sh), 553, 595 (nm).

[0049] Free Base Polymer (Structure 1).

[0050] To a 250 mL heavy walled glass flask was added 1.00g (0.97 mmol) of [5-(4-aminophenyl)-15-bromo-10,20-bis(3,5-dihexylphenyl)porphyrino] zinc (II), 78 mg (0.15 mmol) of bis[(2-diphenylphosphino)phenyl]methane (DPEphos), 442 mg (1.36 mmol) of cesium carbonate and 200 mL of THF. The suspension was flushed with a stream of argon for 15 min, 21 mg (0.01 mmol) of palladium(II) acetate was added and the argon flushing process was continued for a further 10 min. The flask was sealed with a Teflon™ screw plug and the reaction mixture was stirred at 67° C. for 42 h. TLC (methylene chloride/hexanes, 1/1) of the reaction mixture indicated that most or all the starting material had been consumed and that many oligomers had formed. The reaction mixture was filtered through celite and the residual material

where n is an integer.

2. The polymer of claim 1, wherein n is 1 to 5.
3. The polymer of claim 1, wherein n is greater than 5.
4. A device comprising the polymer of claim 1.
5. The device of claim 4, wherein the device is an electronic device.
6. The device of claim 4, wherein the device is an optoelectronic device.
7. The device of claim 4, wherein the device is a photovoltaic device.
8. A method of synthesizing a porphyrin polymer comprising one or more of the reactions shown in FIGS. 2A and 2B.
9. The method of claim 8, wherein the method comprises two or more of the reactions shown in FIGS. 2A and 2B.
10. The method of claim 8, wherein the method comprises three or more of the reactions shown in FIGS. 2A and 2B.
11. The method of claim 8, wherein the method comprises four or more of the reactions shown in FIGS. 2A and 2B.
12. The method of claim 8, wherein the method comprises five or more of the reactions shown in FIGS. 2A and 2B.

13. The method of claim 8, wherein the method comprises six or more of the reactions shown in FIGS. 2A and 2B.

14. The method of claim 8, wherein the method comprises the reactions shown in FIGS. 2A and 2B.

15. A substrate comprising a porphyrin coupled to the surface of the substrate.

16. The substrate of claim 15, wherein the porphyrin is coupled to the surface of the substrate via an aminopropylsiloxy group.

17. The substrate of claim 15, wherein the substrate comprises indium tin oxide.

18. A device comprising the substrate of claim 15.

19. The device of claim 18, wherein the device is a photovoltaic device.

20. The device of claim 18, wherein the device is an electronic or opto-electronic device.

* * * * *

专利名称(译)	可溶性卟啉聚合物		
公开(公告)号	US20130324686A1	公开(公告)日	2013-12-05
申请号	US13/907452	申请日	2013-05-31
[标]申请(专利权)人(译)	GUST JR JOHN 德文斯 利德尔 PAUL ANTHONY		
申请(专利权)人(译)	GUST, JR., JOHN 德文斯 利德尔, PAUL ANTHONY		
当前申请(专利权)人(译)	亚利桑那板校董, 为及代表亚利桑那州立大学亚利桑那州的演技状态的法人团体		
[标]发明人	GUST JR JOHN DEVENS LIDDELL PAUL ANTHONY		
发明人	GUST, JR., JOHN DEVENS LIDDELL, PAUL ANTHONY		
IPC分类号	H01L51/00		
CPC分类号	H01L51/0077 H01L51/0035 C08G73/0273 C08G73/0611 H01L51/009 H01L51/0504 H01L51/42 Y02E10/549		
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其他公开文献	US9076972		
外部链接	Espacenet USPTO		

摘要(译)

结构1的卟啉聚合物, 其中n是整数(例如, 1,2,3,4,5或更大), 通过图1和2中所示的方法合成。图2A和2B。结构1的卟啉聚合物可溶于有机溶剂如2-MeTHF等, 并且可以本体合成(即, 在电聚合以外的方法中)。这些卟啉聚合物具有长激发态寿命, 使得该材料适合作为有机电子器件的有机半导体, 包括晶体管和存储器, 以及太阳能电池, 传感器, 发光器件和其他光电器件。

